

Chapter 08

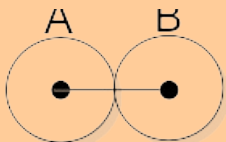
Chemical

Bonding

Bond Parameters:

The equilibrium distance between the two bonded atoms (or) the distance between the two centres of bonded atoms.

- It is generally measured in nanometers (nm) or Angstroms $\text{\AA}$.



Bond energy:

The energy required to break α -bond and obtain constituent atoms is called as bond energy.

It will be measured in KJ/mol (or) Kcal/mol.

Drawbacks of electronic theory:

- It assume that covalent bond is formed due to sharing of pair of electrons.

While sharing identical electrons, they must have same bond length

but molecules are having different length.

- The bond angles (Angles between two bonds) and shapes of molecules are different. This is not been explained by Lewi's theory.

