

Chapter 4

METALS & NON-METALS

Physical properties of Non-metals:

- Non-metals are not malleable and ductile in nature.

They are brittle. When they are hammered with hard tool. They convert into powdered mass.

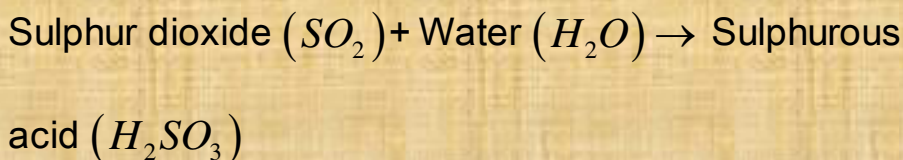
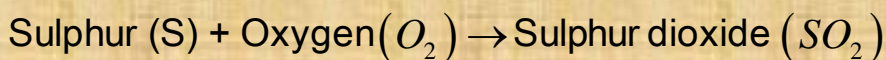
- Non-metals are not sonorous in nature.
- Non-metals have no lustrous nature except Iodine.
- Non-metals are not the good conductors of heat and electricity except graphite (Graphite is the allotrope of non-metal carbon).

Note: Metals like Na (Sodium) and K (Potassium) are soft metals. They can be cut with a knife. Mercury (Hg) is a liquid metal.

Chemical properties of Non-metals:

- Non-metals when react with oxygen forms their oxides.

- Non-metal oxides are acidic in nature. When they dissolved in water forms corresponding acids.



- These acids turns blue litmus paper into red
- Generally non – metals do not react with water

Note:

Highly reactive non – metal phosphorus is stored under water to prevent its reaction with oxygen when exposed to air.

- Non – metals do not reaction with acids.
- Reactions of non-metals with bases are complex in nature.